Low Temperature Symmetry of Pyrochlore Oxide Cd₂Re₂O₇

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We report the X-ray study for the pyrochlore oxide $Cd_2Re_2O_7$. Two symmetry-lowering structural transitions were observed at T_{s1} =200 K and T_{s2} =120 K. The former is of the second order from the ideal cubic pyrochlore structure with space group $Fd\bar{3}m$ to a tetragonally distorted structure with $I\bar{4}m2$, while the latter is of the first order likely to another tetragonal space group $I4_122$. We discuss the feature of the lattice deformation.

A series of pyrochlore oxides $A_2B_2O_7$ show a variety of interesting conductive and magnetic properties, depending on the substitutions of A and B sites [1]. Geometrical frustration on the pyrochlore lattice leads to an exotic ground state such as spin ice for magnetic insulators [2], while for metallic systems a unique heavy carrier behavior is found in spinel LiV_2O_4 having a pyrochlore lattice made of V atoms [3]. In the pyrochlore structure, the A site occupied by a rare-earth or a post-transition metal cation forms a distorted (6+2) coordination, while the B site occupied by a transition metal cation forms a distorted octahedron BO_6 with an equal B-O distance. The structure comprises two interpenetrating pyrochlore lattices composed of the A or B sites (see Fig. 1).

Recently, Hanawa et al. and the others reported the first superconductor $Cd_2Re_2O_7$ ($T_c=1$ K) in the family of pyrochlore oxides [4, 5, 6]. The compound exhibits two more phase transitions at T_{s1} =200 K and T_{s2} =120 K, where electrical and magnetic properties change dramatically [7, 8, 9]. The former is of the second order, while the latter is of the first order. These transitions are accompanied by structural transitions. Three phases appearing are named as phase I $(T>T_{s1})$, II $(T_{s1}>T>T_{s2})$, and III $(T < T_{s2})$. Phase I crystallizes in the ideal pyrochlore structure at room temperature with space group $Fd\bar{3}m$. It was reported in the recent structure study that the crystal system is still cubic within the experimental resolution in the wide temperature range down to T=10 K [8]. Moreover, taking into account the extinction rule of reflections observed below $T_{\rm s1}$ in the single crystal XRD experiments, a possible cubic space group of $F\bar{4}3m$ was suggested. However, recent Re nuclear quadrupole resonance (NQR) experiments have indicated a lack of a threefold axis below $T_{\rm s1}$ implying that the true symmetry is lower than cubic [10]. In addition, very recent highresolution XRD measurements using single crystal barely detected a small splitting of cubic Bragg peaks below T_{s1} , probably due to tetragonal deformation [11]. In this letter, we report the structural study on the low temperature phases of Cd₂Re₂O₇ by means of X-ray diffraction and discuss possible space groups for them.

A single crystal was synthesized as described in the previous paper [4]. We took oscillation photographs using an imaging plate type Weissenberg camera (Mac Sci-

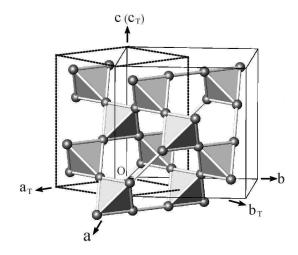


FIG. 1: Pyrochlore lattice on Re (=B site) atoms with cubic and tetragonal cell shown. A body-centered tetragonal lattice made by the transformation $\mathbf{a}_{\mathrm{T}} = (\mathbf{a} - \mathbf{b})/2$, $\mathbf{b}_{\mathrm{T}} = (\mathbf{a} + \mathbf{b})/2$ and $\mathbf{c}_{\mathrm{T}} = \mathbf{c}$ is illustrated by the dashed line.

ence DIP320V) equipped with a closed-cycle helium refrigerator, which is a specialized system with very low background, in the temperature range of 10 K–300 K. The intensity data was collected by a four-circle diffractometer (Mac Science MXC) using a 2θ - ω scan mode and a CCD area detector (Bruker SMART APEX) equipped with a helium flow type cooler in the temperature range of 85 K–300 K. These measurements were performed using a 21 kW rotating-anode X-ray generator with a graphite monochromated Mo- $K\alpha$ radiation. The cooling or heating rate was fixed to 0.5 K/min around the first-order transition at $T_{\rm s2}$.

First, we carefully took oscillation photographs in the whole temperature range. Neither additional reflection breaking the face-centered lattice nor peak splitting of reflections were observed. However, the forbidden reflections were found below $T_{\rm s1}{=}200$ K. Figure 2 (a) shows the temperature dependence of peak intensity for two types of reflections $(0kl:k+l{\neq}4n,00l:l{\neq}4n)$, which are extinguished in $Fd\bar{3}m$. They increase gradually below $T_{\rm s1}$ with decreasing temperature, indicating that the transition is of the second order. Moreover, the slight anomaly was observed for some reflections at $T_{\rm s2}{=}120$ K.

FIG. 3: Group-subgroup relations of symmetry reductions for the maximum subgroups, using the conventional space group setting. Symmetry reductions from face-centered to body-centered lattice are described by the translational-equivalent of index 3 (t3) with the cell transformation, and the other reductions are described by t2.

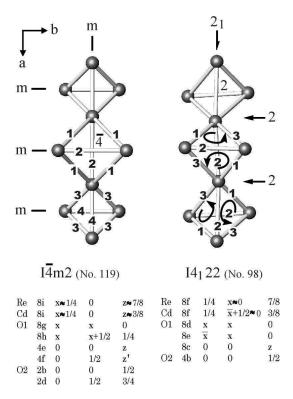


FIG. 4: Schematic representations of Re tetrahedra and atom positions for $I\overline{4}m2$ (left) and $I4_122$ (right). Deformation of Re tetrahedra expected for each space group is illustrated exaggeratedly. Numbers on the bonds for $I\overline{4}m2$ and $I4_122$ correspond to four and three kinds of bond lengths, respectively. Rotational arrows imply a possible bondchirality pattern. Parameter $x,z(\approx x_0-1/8)$ and $z'(\approx x_0+1/8)$ on O1 atom is related to the distortion of the ReO₆ octahedron. The value of x_0 =0.3186(6) is estimated from our structural data at 300 K, which is different from x_0 =0.3089(7) in Ref. 7.

Next, we discuss the symmetry of phase III. We checked all the body-centered tetragonal lattices, and found that only the space groups shown in Fig. 3 could construct the pyrochlore structure. Since the transition at $T_{\rm s2}$ is of the first order, the space group at phase III seems not to be included in the subgroup of $I\bar{4}m2$. Therefore, possible space groups are $I4_1/amd$, $I\overline{4}2d$, $I4_1/a$, $I4_122$, $I4_1md$ and $I4_1$. Here, we can exclude $I4_1/amd$, $I\overline{4}2d$, $I4_1/a$ and $I4_1md$, because these are not compatible with the observed extinction condition of reflections with considering a possible crystal twinning. On the other hand, there is no way to distinguish a true one among the other two space groups of $I4_122$ and $I4_1$ from our experiments. However, the CBED experiments suggested that the point group of phase III is 422 [13]. This fact makes one space group $I4_122$ selected unambiguously. Thus, it is plausible that a symmetry change from $I\bar{4}m2$ to $I4_122$ occurs at T_{s2} , which means missing a mirror symmetry and adding a twofold axis. Such an exchange in symmetry operations is generally seen in a first-order phase transition. It is interesting to note that the Re

atom coordinate changes from $(x\approx 1/4,0,z\approx 7/8)$ in $I\bar{4}m2$ to $(1/4,x\approx0.7/8)$ in $I4_122$. Then, in the $I4_122$ structure, the three kinds of Re-Re bond lengths exist, and the each volume of Re tetrahedon is equivalent. Let's call the three bonds as 1, 2 and 3 in order of bond length. When we trace the bonds of each tetrahedral face as $1\rightarrow2\rightarrow3$, as shown in Fig. 4, a left-rotation (counter-clockwise) configuration can be defined. Then, all the tetrahedral faces possess the same left-rotation configuration. It is naturally expected that a right-rotation configuration is also realized with equal probability. We call this topological chirality on the Re tetrahedra "bondchirality" by analogy with the R-S convention for a chiral molecule in organic chemistry. The direction of rotation depends on the sign of the Re atom parameter. In a real crystal, two types of domain with the left and right rotations are formed below $T_{\rm s2}$ as a racemic mixture. The Re tetrahedron in $I\overline{4}m2$ is considered to be a disordered state without bondchirality. This means that the transition at $T_{\rm s2}$ is relevant to ordering of the bondchirality.

In summary, we have reported the low temperature X-ray study for the first pyrochlore oxide superconductor $\mathrm{Cd_2Re_2O_7}$. The breaking of the cubic lattice and inversion symmetry are observed below $T_{\mathrm{s1}}{=}200~\mathrm{K}$ as well as below $T_{\mathrm{s2}}{=}120~\mathrm{K}$. From the experimental findings and the group-subgroup relations, it is suggested that the space group changes from $Fd\bar{3}m$ to $I\bar{4}m2$ at T_{s1} , and to $I4_122$ at T_{s2} . In addition, the characteristic deformations of Re tetrahedron were discussed in each phase.

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M. A. Subramanian, G. Aravamudan and G. V. Subba Rao, Prog. Solid State Chem. 15, 55 (1983).

^[2] S. T. Bramwell and M. J. P. Gingras, Science **294**, 1495 (2001)

^[3] C. Urano, M. Nohara, S. Kondo, F. Sakai, H. Takagi, T. Shiraki and T. Okubo, Phys. Rev. Lett. 85, 1052 (2000).

^[4] M. Hanawa, Y. Muraoka, T. Tayama, T. Sakakibara, J. Yamaura and Z. Hiroi, Phys. Rev. Lett. 87, 187001 (2001).

^[5] H. Sakai, K. Yoshimura, H. Ohno, H. Kato, S. Kambe, R. E. Walstedt, T. D. Matsuda, Y. Haga and Y. Onuki, J. Phys. Condens. Matter 13, L785 (2001).

^[6] R. Jin, J. He, S. McCall, C. S. Alexander, F. Drymiotis and D. Mandrus, Phys. Rev. B 64, 180503 (2001).

^[7] P. C. Donohue, J. M. Longo, R. D. Rosenstein and L. Katz, Inorg. Chem. 4, 1152 (1965).

^[8] M. Hanawa, J. Yamaura, Y. Muraoka, F. Sakai and Z. Hiroi, J. Phys. Chem. Solids 63, 1027 (2002).

^[9] Z. Hiroi, J. Yamaura, Y. Muraoka and M. Hanawa, J.

- Phys. Soc. Jpn. **71**, 1634 (2002).
- [10] O. Vyaselev, K. Arai, K. Kobayashi, J. Yamazaki, K. Kodama, M. Takigawa, M. Hanawa and Z. Hiroi, Phys. Rev. Lett. 89, 017001 (2002).
- [11] J. P. Castellan, B. D. Gaulin, J. van Duijn, M. J. Lewis, M. D. Lumsden, R. Jin, J. He, S. E. Nagler and D. Mandrus, cond-mat/0201513.
- [12] International Tables for Crystallography, ed. T. Hahn (Kluwer Academic Publishers, Dordrecht, 1989) vol. A, 2nd edn.
- [13] K. Tsuda, M. Oishi, M. Tanaka, M. Hanawa and Z. Hiroi, Meeting Abstracts of the Phys. Soc. Jpn. (Annual Meeting, March, 2002) Pt. 3, p. 562.